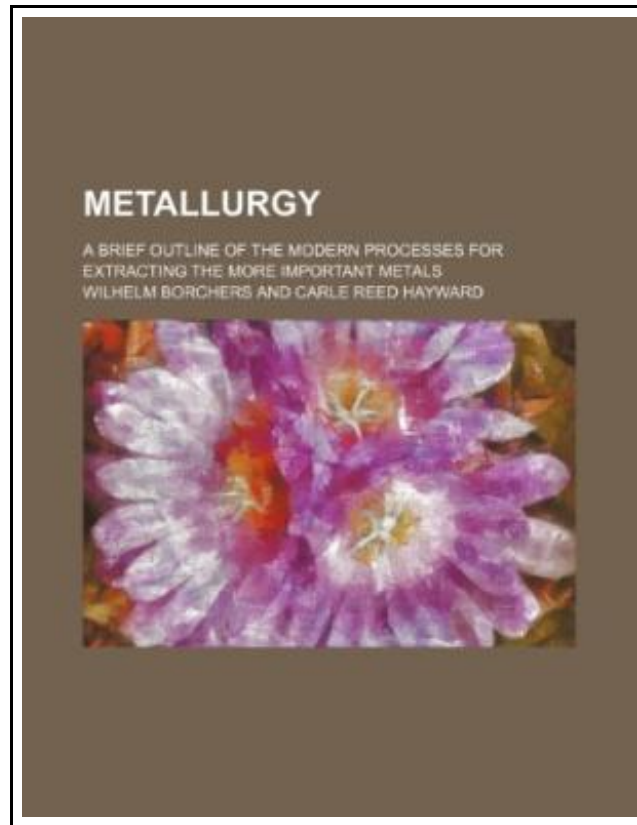


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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1911 Excerpt: .carbon. This reduction of the NiO was first accomplished by the method of Fleitmann, who fused with Mg, which is an extremely energetic reducing agent; then Basse and Selve accomplished the same end by mixing in as much as 3 of MnO before carrying out the reducing roast. The MnO₂ is reduced, together with the NiO, and serves to remove the NiO when the cube nickel is heated in carbon-free crucibles; naturally it prevents the formation of NiO by atmospheric oxygen when the metal is melted. Electrolysis is employed if the nickel contains precious metals. Since the solution tension of Ni is on the other side of H, the concentration of H ions must be kept as low as possible, contrary to the practice in the electrolytic refining of Cu. This is accomplished by limiting the amount of free acids having high dissociation constants (HCl, H₂SO₄) or by using acids with low dissociation constants (H₃BO₃, H₃PO₄, and organic acids). Unfavorable changes in concentration are avoided by moving the electrodes and heating the electrolyte. When these requirements are fulfilled, the electrolytic deposition of Ni offers no difficulties worth mentioning. The conditions for carrying out the electrolysis are as follows: Electrolyzing Tanks: made of wood with lead lining. Anodes: Cast plates of impure nickel, containing precious metals with S up to 3 . Cathodes: Pure Ni. Electrolyte: NiSO₄ or NiCl₂ solution, the former being preferred. Concentration 4 to 14 oz. Ni to the gallon not under 0.01 or...

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