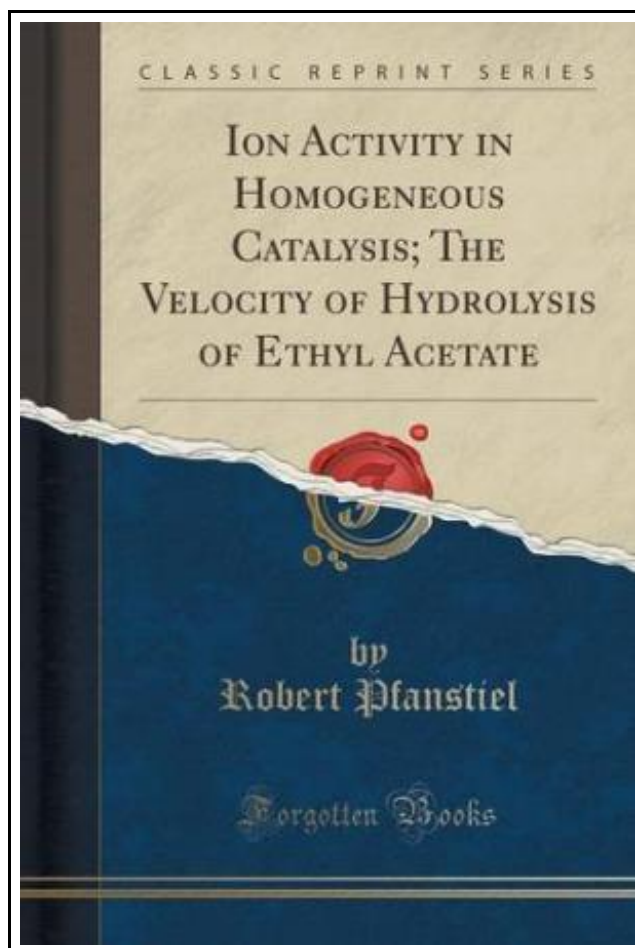


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Forgotten Books, United States, 2015. Paperback. Book Condition: New. 229 x 152 mm. Language: English . Brand New Book ***** Print on Demand *****.Excerpt from Ion Activity in Homogeneous Catalysis; The Velocity of Hydrolysis of Ethyl Acetate As a result of considerable evidence, Mac Innes (Jour. Amer. Chem. Soc. 41, 1086 [1919]) has arrived at the conclusion that in solutions of the same molality of hydrochloric acid, lithium, sodium, and potassium chlorides, the chlorine ion has the same activity. He further made the assumption that in a solution of a given strength, the activities of the potassium and chlorine ions are the same. These hypotheses received considerable confirmation in dilute solutions from the electromotive force measurements of Ming Chow (Jour. Amer. Chem. Soc. 42, 477 [1920]) and in concentrated solutions by Harned (Jour. Amer. Chem. Soc. 42, 1808 [1920]). On the basis of these assumptions, Harned calculated from existing electromotive force data the individual ion activity coefficients of these uni-univalent electrolytes. If Mac Innes assumptions are correct it follows from these calculations that the activity coefficient of the hydrogen ion in dilute solutions of hydrochloric acid decreases until a concentration of 0.15 M. is reached and then increases quite rapidly. In any event, this activity coefficient must exhibit a minimum in the neighborhood of from 0.1 M. to 0.2 M. concentration. These conclusions, if true, or valid within narrow limits, will be of considerable importance in the calculation of equilibria in solutions as well as homogeneous catalysis. Consequently, this investigation was undertaken with the purpose of finding out whether further support for the above hypothesis could be obtained from a study of hydrogen ion catalysis. It has been found that the monomolecular velocity constant of hydrolysis of ethyl acetate in dilute solutions of hydrochloric acid is roughly proportional to the...



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